

I'm not robot!

4 EASY STEPS TO BALANCING CHEMICAL EQUATIONS

STEP 1: List down atoms

$$\text{C}_2\text{H}_5\text{F} + \text{O}_2 \rightarrow \text{CO} + \text{H}_2\text{O} + \text{F}_2$$

C
H
F
O

STEP 2: Count atoms on both sides

$$\text{C}_2\text{H}_5\text{F} + \text{O}_2 \rightarrow \text{CO} + \text{H}_2\text{O} + \text{F}_2$$

6 C 1
5 H 2
1 F 2
2 O 2

STEP 3: Add big numbers in front

$$2\text{C}_2\text{H}_5\text{F} + 7\text{O}_2 \rightarrow 6\text{CO} + 5\text{H}_2\text{O} + \text{F}_2$$

12 C + 5 12
10 H 2 10
2 + F 2
17 2 O 2 + 5 17

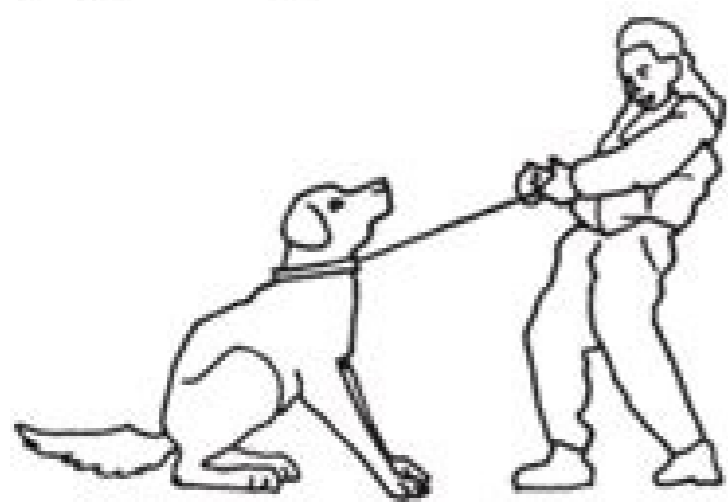
STEP 4: Multiply across (if necessary)

$$[2\text{C}_2\text{H}_5\text{F} + 7\text{O}_2 \rightarrow 6\text{CO} + 5\text{H}_2\text{O} + \text{F}_2] \times 2$$

$$4\text{C}_2\text{H}_5\text{F} + 14\text{O}_2 \rightarrow 12\text{CO} + 10\text{H}_2\text{O} + 2\text{F}_2$$

Rules for word equations	Reactants	Products
1. Write the names of the reactants.		
2. Write the names of the products.		
3. Balance the equation.		
4. Check the balanced equation.		

- c If the forward and backward forces are balanced, what will happen to the speed of the car?
..... (1)
- d If the forward force is bigger than the backward force, what will happen to the car?
..... (1)
- e If the backward force is bigger than the forward force, what will happen to the car?
..... (1)
- f Give the name of another force which could be acting on the car.
..... (1)
- g Which direction will it be acting in?
..... (1)
- 4 Linda is trying to get the dog to move!



- a Draw an arrow on the diagram to show the direction of the force Linda is applying. (1)
- b The dog is pulling on the rope. Draw an arrow on the diagram to show direction of this force if the dog does not move. (1)
- c When the rope breaks, what will happen to:
- i the dog?
..... (1)
- ii Linda?
..... (1)

Total marks = 32

Rules for Balancing Chemical Equations
1. Adjust the coefficients not the subscripts.
2. Start with the most complex molecule.
3. Balance polyatomic ions as a whole.
4. Check the balanced equation.

Patterns of behaviour – 1: Reactions of metals

- 1a i Magnesium burns very brightly when heated in air. Complete the word equation:
Magnesium + → (2)
- ii What **type** of reaction is this? (1)
- b Three gas jars are shown below. Each contains a different gas and a piece of burning magnesium is placed into each and observed:
- A

Exhaled air

B

Nitrogen

C

Oxygen
- i In which gas jar will the magnesium burn **even more** brightly than it did in air?
..... (1)
- ii Explain your answer.
..... (1)
- iii In which gas jar will it **go out** immediately? (1)
- iv Explain your answer.
..... (2)
- c i Name **one** metal which does not react with oxygen even when heated.
..... (1)
- ii What is this metal often used for?
..... (1)
- iii Explain why it is used for this purpose.
..... (1)

Chemistry is a conceptual subject and, in order to explain many of these concepts, teachers use models to describe and explain the microscopic world and relate it to the macroscopic properties of matter. This resource is designed to provide strategies for dealing with some of the misconceptions that students have in the form of ready-to-use classroom resources. Last updated2 October 2012Hopefully a phased worksheet that aims to gradually improved students confidence with word and symbol equations.Tes classic free licenceSelect overall rating(no rating>Your rating is required to reflect your happiness.Write a reviewUpdate existing reviewIt's good to leave some feedback.Something went wrong, please try again later.Thanks for sharingEmpty reply does not make any sense for the end userThank you very much. This is a great resource. Will be using it for my after school with some year 9 students. Empty reply does not make any sense for the end userPerfect thanks, Year 8 loved this as part of their 'Speedy Chemistry'Empty reply does not make any sense for the end userEmpty reply does not make any sense for the end userThanks. I used the worksheet as a starter towards revision.Empty reply does not make any sense for the end userReport this resourceto let us know if it violates our terms and conditions. Our customer service team will review your report and will be in touch. Chemical reactions happen all around us. When we burn methane in a Bunsen burner, when we cook at home, and even inside our bodies, elements and compounds are reacting together. Elements are made of only one type of atom. And compounds are made from different types of atoms chemically bonded together. Reactions happen when bonds break and new bonds are made between different atoms. To help understand what's going on, it's useful to represent the reaction with a word equation. Let me show you one of my favourite reactions. This fine powder is iron. When we sprinkle it over the flame, it lights up like a sparkler. The iron is reacting with the oxygen in the air to make a new compound, called iron oxide. From that information, we can write a word equation to represent the reaction. The reactants are iron and oxygen, they go on the left hand side. Iron plus oxygen. We use an arrow to show the direction of the reaction. You can think of the arrow as saying 'reacts to form'. So, iron plus oxygen reacts to form...The last part is the product, iron oxide, which we write on the right hand side. Iron plus oxygen reacts to form iron oxide. And that's your word equation!

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